\mathcal{A}

configuration. This would imply that for the *S* ligand, Co-A and Cr-A have the same configuration and also Co-B and Cr-B have identical configurations. An analogous relationship exists between Co-A' and Cr-A' and between $Co-B'$ and $Cr-B'$, where the ligand has the R configuration.

Although the ligand field absorption band of the chromium(III) complex is \sim 1500 cm⁻¹ toward higher energy than that of the cobalt(II1) complex, isomers B and \bar{B}' for both metals have larger ϵ_M values than isomers A and A'. Co-A and Cr-A both show positive CD's at the low-energy end of the absorption; Co-B and Cr-B both show stronger, negative CD's at this energy. Similarly, Co-A' and Cr-A' both have negative CD's whereas Co-B' and Cr-B' show stronger, positive CD's. Other pairs of tris cobalt(II1) and chromium(II1) complexes for which isomers believed to have the same absolute configuration show similar CD curves (the transition for chromium(III) occurring at higher energy in each case) include those of ethylenediamine, 1,2-propanediamine, 6 and acetylcamphor.⁷ An exception to this behavior has been found for the tris-biguanide complexes.*

The above arguments strongly suggest but do not prove that the chromium(II1) and cobalt(II1) diastereomers showing parallel chromatographic behavior have identical absolute configurations. The B and B' isomers were shown' to be thermodynamically more stable than A and A' for cobalt(III) . However the chromium(II1) isomers for a given ligand are formed in essentially the statistical 1 : 1 ratio, and no isomerization can be detected by tlc after refluxing chlaroform solutions of the isomers for 24 hr in the presence of charcoal.

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The Crystal Structure of High-pressure Synthesized Holmium Diantimonide

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The synthesis of the rare earth diantimonides, with the exception of $EuSb₂$, has been previously reported.¹⁻³ Up to and including $SmSb₂$, the synthesis is conventional; the structure is of the LaSb_2 type.¹ For Gd and Tb, this structure type, along with a new structure type, is synthesized *via* high pressure.2 This second

(1) R. Wang and H. Steinfink, *Inovg. Chem.,* **6,** 1685 (1967).

type, which has been found for Gd through Lu with the exception of Yb, has been reported to be orthorhombic with two molecules per unit cell.^{2,3}

In an attempt to determine, by single-crystal methods, the structure of this apparently new type, we have prepared HoSb₂ using high-pressure techniques.

Experimental Section

Holmium, filed in an inert atmosphere, was mixed with antimony in a 1:2 gram-atom ratio. This mixture was then subjected to 65 kbars and 1000° for 5 hr in a girdle-anvil device. Powder patterns were compared with the published LuSb₂ pattern to establish whether the synthesis was successful. Not all of the lines of LuSb₂ were present in patterns for several preparations of HoSb₂. Debye-Scherrer techniques (λ (Cu K α_1) 1.54050 **A)** were used throughout this investigation because the expected single crystals did not materialize.

Results **and** Discussion

The orthorhombic indexing of Eatough and Hall was suspected to be incorrect. As one piece of evidence, the $d_{\text{o}}-d_{\text{e}}$ comparison is unsatisfactory. Secondly, it was noted that the published cell parameter *a* could not be obtained from the reported pattern and indexing. It is in error by 0.1 **A.** Finally, while some of our patterns exhibited the line corresponding to the first line of the L u Sb_2 pattern, it was missing in others. Suspecting the LuSb₂ pattern to be a two-phase pattern, we have worked with patterns which did not show this line. A satisfactory hand indexing was accomplished for a hexagonal cell with $a = 3.33$ Å and $c = 7.83$ Å. Since this volume is sufficient for only one molecule of

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HoSbz, it is an easy matter exhaustively to check all possible structures. No ordered structure which gave reasonable intensity agreement was found for this cell. A disordered model which gave good agreement led to an ordered structure in a larger pseudoorthohexagonal cell. The unit cell parameters are $a = 3.343 \pm 0.001$ A, $b = \sqrt{3}a = 5.790 \pm 0.001$ A, and $c = 7.840 \pm 0.001$ 0.002 Å with $Z = 2$ and a density of 8.94 g/cm³. Atomic positions, using space group C222, are as follows: Ho, 2a $(0, 0, 0)$; Sb, 4k $(\frac{1}{4}, \frac{1}{4}, 0.34)$. The value of *z* for Sb was chosen to make $d_{\text{Ho-}Sb}/(r_{\text{Ho}} +$ $r_{\rm Sb}$) = $d_{\rm Sb-Sb}/2r_{\rm Sb}$ using the radii of Teatum, Gschneidner, and Waber.⁴

The calculated powder pattern, using a cylindrical absorption correction $(\mu r = 19.2)$, temperature factors of $B = 1.0$ Å², and the program of Smith,^{δ} is listed together with the observed pattern in Table I. The structure of $HoSb₂$ is shown in Figure 1; distances are given in Table 11.

TABLE **I1**

We are well aware of the fact that the structure presented may indeed be only a subcell of the true structure. As stated earlier, single crystals were not obtained; thus we must infer, in view of the excellent intensity fit, that this structure is at least qualitatively correct and definitely represents an improvement over the earlier work. Unit cell constants for $Gd-TmSb₂$ cannot be determined by us because no data for these have been published. Cell constants have been determined for $L \text{uSb}_2$ using the published pattern, this structure, and a least-squares program. The values are $a = 3.3211 \pm 0.0005 \text{ Å}, b = 5.7521 \pm 0.0009 \text{ Å}, \text{and}$ $c = 7.782 \pm 0.001$ Å. The calculated density is 9.35 $g/cm³$. It is noted that these cell constants are in disagreement with those of Eatough and Hall by less than

Figure 1.—Crystal structure of H_0Sb_2 .

 3% . These differences are, however, from 12 to 19 times the standard errors given by them.

This structure accurately accounts for the published pattern of $LuSb₂$ in regard to intensities as well as d spacings. Two lines are, however, not accounted for by this structure; they are at $d_1 = 3.09$ and $d_2 = 1.551$. \Ye have not established what these lines come from but note that they probably belong to the same impurity since $d_1 = 2 \cdot d_2$.

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