configuration. This would imply that for the S ligand, Co-A and Cr-A have the same configuration and also Co-B and Cr-B have identical configurations. An analogous relationship exists between Co-A' and Cr-A' and between Co-B' and Cr-B', where the ligand has the Rconfiguration.

Although the ligand field absorption band of the chromium(III) complex is ~ 1500 cm⁻¹ toward higher energy than that of the cobalt(III) complex, isomers B and B' for both metals have larger ϵ_M values than isomers A and A'. Co-A and Cr-A both show positive CD's at the low-energy end of the absorption; Co-B and Cr-B both show stronger, negative CD's at this energy. Similarly, Co-A' and Cr-A' both have negative CD's whereas Co-B' and Cr-B' show stronger, positive CD's. Other pairs of tris cobalt(III) and chromium(III) complexes for which isomers believed to have the same absolute configuration show similar CD curves (the transition for chromium(III) occurring at higher energy in each case) include those of ethylenediamine, 1,2-propanediamine,⁶ and acetylcamphor.⁷ An exception to this behavior has been found for the tris-biguanide complexes.8

The above arguments strongly suggest but do not prove that the chromium(III) and cobalt(III) diastereomers showing parallel chromatographic behavior have identical absolute configurations. The B and B' isomers were shown¹ to be thermodynamically more stable than A and A' for cobalt(III). However the chromium(III) isomers for a given ligand are formed in essentially the statistical 1:1 ratio, and no isomerization can be detected by the after refluxing chloroform solutions of the isomers for 24 hr in the presence of charcoal.

Acknowledgment.—Support for this research has been provided by the University of Kansas General Research Fund and by the National Science Foundation under Grant GP-19877.

(6) (a) A. J. McCaffery, S. F. Mason, and R. E. Ballard, J. Chem. Soc. A, 2883 (1965); (b) A. J. McCaffery, S. F. Mason, and B. J. Norman, Chem. Commun., 49 (1965).

(7) (a) R. M. King and G. W. Everett, Jr., Inorg. Chem., 10, 1237 (1971);
(b) W. DeW. Horrocks, Jr., D. L. Johnston, and D. MacInnes, J. Amer. Chem. Soc., 92, 7620 (1970).

(8) G. R. Brubaker and L. E. Webb, ibid., 91, 7199 (1969).

CONTRIBUTION FROM THE LAWRENCE RADIATION LABORATORY, UNIVERSITY OF CALIFORNIA, LIVERMORE, CALIFORNIA 94550

The Crystal Structure of High-Pressure Synthesized Holmium Diantimonide

By QUINTIN JOHNSON

Received February 5, 1971

The synthesis of the rare earth diantimonides, with the exception of EuSb₂, has been previously reported.¹⁻³ Up to and including SmSb₂, the synthesis is conventional; the structure is of the LaSb₂ type.¹ For Gd and Tb, this structure type, along with a new structure type, is synthesized *via* high pressure.² This second

(1) R. Wang and H. Steinfink, Inorg. Chem., 6, 1685 (1967).

TABLE I				
OBSERVED AND	CALCULATED	Powder	Pattern	of HoSb ₂

Pseudo-	
hever-	Orth

hexag-	Ortho-			-	
indices	indices	Caled	Obsd	Caled	Obed
001	001	7 840		1	0.054
002	002	3 920		1	• • •
100	110	2 895	2 80	15	
200	020	2.000	2.00	10	vv
101	111	2.716	2.709	100	VS
	021		2.100	100	• 3
003	003	2.613	2.608	35	m
102	112	2.329	2.327	81	s
	022			01	^D
004	004	1.960		0	
103	113	1.940	1.939	17	w
	023				
110	200	1.671	1.669	7	vw
	130				
111	201	1.635	1.634	48	m
	131				
104	114	1.623	1.624	48	m
	024				
005	005	1.568		1	
112	202	1.538	1.539	42	m
	132				
200	220	1.448	1.447	43	m
	040				
201	221	1.423	• • •	1	
	041				
113	203	1.408	1.405	10	vw
105	133	1 0 70	1 0 7 0		
105	115	1,378	1.378	34	m
000	025	1.0.70		0	
202	222	1.358	• • •	2	• • •
006	042	1 907	1 207	10	
114	000	1.307	1.307	12	w
114	204	1.272	1.270	30	m
002	104	1 966	1 966	65	
200	223 043	1.200	1,200	00	m+
	010				

type, which has been found for Gd through Lu with the exception of Yb, has been reported to be orthorhombic with two molecules per unit cell.^{2,3}

In an attempt to determine, by single-crystal methods, the structure of this apparently new type, we have prepared $HoSb_2$ using high-pressure techniques.

Experimental Section

Holmium, filed in an inert atmosphere, was mixed with antimony in a 1:2 gram-atom ratio. This mixture was then subjected to 65 kbars and 1000° for 5 hr in a girdle-anvil device. Powder patterns were compared with the published LuSb₂ pattern to establish whether the synthesis was successful. Not all of the lines of LuSb₂ were present in patterns for several preparations of HoSb₂. Debye–Scherrer techniques (λ (Cu K α_1) 1.54050 Å) were used throughout this investigation because the expected single crystals did not materialize.

Results and Discussion

The orthorhombic indexing of Eatough and Hall was suspected to be incorrect. As one piece of evidence, the d_o - d_o comparison is unsatisfactory. Secondly, it was noted that the published cell parameter a could not be obtained from the reported pattern and indexing. It is in error by 0.1 Å. Finally, while some of our patterns exhibited the line corresponding to the first line of the LuSb₂ pattern, it was missing in others. Suspecting the LuSb₂ pattern to be a two-phase pattern, we have worked with patterns which did not show this line. A satisfactory hand indexing was accomplished for a hexagonal cell with a = 3.33 Å and c = 7.83 Å. Since this volume is sufficient for only one molecule of

 ⁽²⁾ N. L. Eatough and H. T. Hall, *ibid.*, 8, 1439 (1969).

⁽³⁾ N. L. Eatough and H. T. Hall, ibid., 9, 416 (1970).

HoSb₂, it is an easy matter exhaustively to check all possible structures. No ordered structure which gave reasonable intensity agreement was found for this cell. A disordered model which gave good agreement led to an ordered structure in a larger pseudoorthohexagonal cell. The unit cell parameters are $a = 3.343 \pm 0.001$ Å, $b = \sqrt{3a} = 5.790 \pm 0.001$ Å, and $c = 7.840 \pm 0.002$ Å with Z = 2 and a density of 8.94 g/cm³. Atomic positions, using space group C222, are as follows: Ho, 2a (0, 0, 0); Sb, 4k $(^{1}/_{4}, ^{1}/_{4}, 0.34)$. The value of z for Sb was chosen to make $d_{\rm Ho-Sb}/(r_{\rm Ho} + r_{\rm Sb}) = d_{\rm Sb-Sb}/2r_{\rm Sb}$ using the radii of Teatum, Gschneidner, and Waber.⁴

The calculated powder pattern, using a cylindrical absorption correction ($\mu r = 19.2$), temperature factors of B = 1.0 Å², and the program of Smith,⁵ is listed together with the observed pattern in Table I. The structure of HoSb₂ is shown in Figure 1; distances are given in Table II.

TABLE II

	Γ	ISTANCES I	N HoSb ₂ (A	Å)	
Ho	4 Sb	3.15	Sb	2 Sb	3.02
	6 Ho	3.34		2 Ho	3.15
				6 Sb	3.34

We are well aware of the fact that the structure presented may indeed be only a subcell of the true structure. As stated earlier, single crystals were not obtained; thus we must infer, in view of the excellent intensity fit, that this structure is at least qualitatively correct and definitely represents an improvement over the earlier work. Unit cell constants for Gd-TmSb₂ cannot be determined by us because no data for these have been published. Cell constants have been determined for LuSb₂ using the published pattern, this structure, and a least-squares program. The values are $a = 3.3211 \pm 0.0005$ Å, $b = 5.7521 \pm 0.0009$ Å, and $c = 7.782 \pm 0.001$ Å. The calculated density is 9.35 g/cm³. It is noted that these cell constants are in disagreement with those of Eatough and Hall by less than



Figure 1.—Crystal structure of HoSb₂.

3%. These differences are, however, from 12 to 19 times the standard errors given by them.

This structure accurately accounts for the published pattern of LuSb₂ in regard to intensities as well as dspacings. Two lines are, however, not accounted for by this structure; they are at $d_1 = 3.09$ and $d_2 = 1.551$. We have not established what these lines come from but note that they probably belong to the same impurity since $d_1 = 2^*d_2$.

Acknowledgment.—We thank Mr. H. Stromberg for the high-pressure synthesis and Mr. V. Silveira for powder photography. This work was performed under the auspices of the U. S. Atomic Energy Commission.

⁽⁴⁾ E. T. Teatum, K. A. Gschneidner, Jr., and J. T. Waber, Report LA-4003, Los Alamos Scientific Laboratory, Los Alamos, N. M., 1968.

⁽⁵⁾ D. K. Smith, Report UCRL-50264, Lawrence Radiation Laboratory, Livermore, Calif., 1967.